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Chapter 1

Effects of Grain Refining on Columnar-to-Equiaxed Transition in Aluminum Alloys

Hicham Tahiri, Serageldin S. Mohamed, Herbert W. Doty, Salvador Valtierra and Fawzy H. Samuel

Additional information is available at the end of the chapter

http://dx.doi.org/10.5772/intechopen.70346

Abstract

The effects of grain refining in ultra-pure aluminum, commercially pure aluminum (1050), and Al-7%Si binary alloy were investigated, using different additions of Al-10%Ti, Al-5%Ti-1%B, and Al-4%B master alloys. Thermal analysis and metallography were used to assess the variations in microstructure resulting from these additions, at solidification rates of 0.8°C/s and ~10°C/s. The results revealed that addition of Al-4%B to ultra-pure aluminum forms AlB$_3$ and AlB$_2$ which have no grain-refining effect. Without grain refiner addition, the pure aluminum microstructure exhibits a mixture of columnar and equiaxed grains. Addition of 30ppm Ti is sufficient to promote equiaxed grains at ~10°C/s but requires addition of 1000 ppm B to obtain similar results at 0.8°C/s. Increasing the Si content to 7% reduces the initial grain size of pure aluminum from 2800 μm to ~1850 μm, and further to 450 μm with addition of ~500ppm B. In commercial aluminum, the B reacts with traces of Ti forming Al$_3$Ti and TiB$_2$ phases which are active grain-refiners. In Al-7%Si, Ti reacts with Si forming (Al,Si)$_2$Ti phase, which is a poor refining agent. This phenomenon is termed poisoning. No interaction between B and Si is observed in the commercial aluminum or Al-7%Si alloy when B is added.

Keywords: aluminum, grain refining, poisoning, columnar-to-equiaxed transition, solidification rate, macrostructure

1. Introduction

Master alloys of the type Al-B are largely used in production of ultra-pure aluminum to react with transition elements such as V, Cr, and Zr [1]. Boron is not considered as
effective a grain refiner when added to pure aluminum [2]. Once Si is added to Al, grain refining is activated together with a change in the α-Al dendritic structure [3]. Two types of B-compounds exist in the Al-B master alloys: AlB$_2$ and AlB$_{12}$ [4]. The AlB$_2$ compound is stable at room temperature and contains 44.5% B [5–7]. It is inferred from the Al-B binary diagram presented in Figure 1(a) that there is a peritectic reaction at 975°C: L + AlB$_{12}$ → AlB$_2$. A eutectic reaction (liquid → α-Al + AlB$_2$) takes place at 660°C, making the maximum solubility of B in Al about 20 ppm where the α-Al grains would precipitate on the AlB$_2$ particles. The melting points of AlB$_2$ and AlB$_{12}$ are 1665 ± 50°C and 2163 ± 50°C, respectively, resulting in the formation of solid dispersoid particles in the molten liquid [8]. Fundamentally speaking, the addition of a grain refiner to the molten metal would result in the nucleation of new grains or reaction with other elements in the molten

![Figure 1](image1.png)

**Figure 1.** (a) Aluminum rich corner of Al-B phase diagram [7]. Broken arrow points to Al + AlB$_2$ eutectic reaction. (b) Schematic drawing showing the graphite mold used for thermal analysis. (c) Casting made in steel cups to obtain a solidification rate of 10°C/s.
metal to form nucleation sites. When Al-4%B master alloy is added to commercial alu-
mimum containing traces of Ti, the B atoms would react with Ti forming TiB\(_2\) phase
\((Ti + B \rightarrow TiB\(_2\)(s)))\). The free energy \((\Delta G)\) associated with this reaction can be calculated as
\(\Delta G = -73381 + 38.996T\), where \(T\) is the liquid temperature [9]. The enthalpy energy associ-
ated with the formation of TiB\(_2\) phase particles is fairly large – 326.41 kJ/mol—which allows
for their stability in the molten metal.

In order to clarify the role of Si in activating the grain refiner, several castings were made
using pure Al (99.999%), commercially pure aluminum (1050), and a binary Al-7%Si alloy.
The study was conducted using thermal analysis, whereas another set of castings for grain
refining was carried out using reduced pressure testing steel cups.

## 2. Experimental procedure

In this study, pure aluminum (99.999%), commercial pure aluminum (1050), and Al-7%Si
were used. The compositions of the materials used are shown in Table 1. Each base alloy
was melted in a resistance-heated clay graphite crucible and held at 750°C. The three grain
refiners Al-10%Ti, Al-5%Ti-1%B, and Al-4%B were added to each melt in the amounts (Ti
or B): 0.02–0.4% wt.%. Prior to casting, the liquid metal was held at 750°C for 10, 30, 60,
90, and 120 min after the grain refiner addition was made and then cast. During the hold-
ing time, the melt was continuously stirred using a graphite impeller to minimize sedimen-
tation of Ti-containing particles and maintain uniformity. Samplings from the different
melts prepared at 750°C were poured into a preheated (600°C) cylindrical graphite mold
to achieve near-equilibrium solidification conditions (solidification rate ~0.8°C/s), as shown
in Figure 1(b). Another set of castings was made using steel cups of the reduced pressure
testing machine in air. The steel cups were not preheated in order to obtain a high cooling
rate of about 10°C/s, as shown in Figure 1(c)—solidification rate is about 10°C/s. Castings
were also made using steel cups preheated at 450°C to represent an intermediate solidifica-
tion rate. Additionally, the microstructure was evaluated at both cooling rates.

The temperature-time data was obtained using a K-type (chromel-alumel) thermocou-
ple inserted through the bottom of the graphite mold along the centerline, its tip reaching
up to about one-third the height of the mold from the bottom. The cooling curves were
recorded using a data acquisition system attached to the thermocouple. Chemical analysis

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Table 1. Chemical composition (wt.%).
of all melts/castings was determined using a Spectrolab Jr CCD Spark Analyzer (average of three burns per sample).

The solidified castings were sectioned perpendicular to the centerline axis of the cylinder, at the level of the thermocouple tip, and polished for metallographic examination, using standard polishing procedures. The polished samples were examined using an optical microscope and an electron probe micro-analyzer (JEOL JXA-8900L operating at 20 kV), equipped with energy dispersive X-ray spectroscopic (EDS) and wavelength dispersive spectroscopic (WDS) facilities.

3. Results and discussion

Figure 2 shows the temperature-time curves of pure Al and Al-7%Si binary alloy after different melt treatments. Figure 2(a) shows that the addition of B to pure Al reduces the undercooling by approximately -1°C, indicating the relative ineffectiveness of B as a grain refiner in the absence of Si. A 300 ppm level of B added to the Al-7%Si alloy eliminated the phenomenon of recalescence. Increasing the concentration of B to 800 ppm has minimal further effect on the alloy solidification behavior. As shown below, the transformation of AlB$_{12}$ to a simple phase AlB$_2$ is the principal parameter in the disappearance of the observed undercooling shown in Figure 2(a).

Figure 2(b) demonstrates the variation in aluminum grain size as a function of Sr-B interaction in the presence of 200 ppm Sr as well as holding time. In the absence of melt treatment, the average grain size was about 2900 μm. Once the two agents (B and Sr) were simultaneously added to the melt, the grain size dropped to 420 μm at 0.1%B. The observed reduction in the grain size may be interpreted in part due to presence of traces of Ti either in the used aluminum or in the added Al-4%B master alloy as shown in Table 2. At higher B concentrations, the grain size tended to increase reaching 3300 μm at holding time of 120 min. This observation may be explained in terms of agglomeration of the nucleation sites [10]. As shown in Table 1, a eutectic reaction takes place at about 660°C; Liquid $\rightarrow$ AlB$_2$ + α (solid). In Si-containing aluminum alloys, i.e., 356 alloy, the melting point is less than the eutectic temperature, which facilitates the formation of AlB$_2$, an active nucleation agent. The melting temperature of pure Al is about 660.5°C, and hence the reported sluggishness of B as a grain refiner in pure Al [11]. Figure 2(c) displays the variation in the grain size of Al-7%Si alloy as a function of holding time when the alloy was grain refined using Al-4%B master alloy in the presence of 200 ppm Sr. Due to high Si content, the initial grain size dropped to approximately 1850 μm. With the addition of approximately 500 ppm B, the grain size was reduced to about 450 μm (~75%). Holding time up to 120 min seems to have a marginal effect on further decrease in the alloy grain size. These findings are in good agreement with the abovementioned discussion.

Figure 3 reveals the possibility of the coexistence of the two B-based compounds in pure aluminum. The composition of these two phases was confirmed using the WDS technique as shown in Table 3. Figure 4 shows the distribution of B- (black), and Si-phases (light gray).
in Al-7%Si alloy. As can be seen, these elements tend to precipitate away from each other (marked zones A and B in Figure 4). This observation is confirmed by the X-ray elemental distribution for each zone shown in Figure 5, indicating that there is no affinity for reaction
between these two elements. In contrast, when Ti is added with B to the 1050 commercial aluminum, a clear interaction between Ti and B takes place leading to the formation of TiB$_2$, as demonstrated in Figure 6. Thus, B has no grain-refining capacity when added as a single element to ultra-pure aluminum, whereas when the metal contains traces of Ti, adding B leads to the possible formation of different nucleation sites that coexist, such as AlB$_{12}$, AlB$_2$, and TiB$_2$ in Al-Si alloys.

It is well established that increasing solidification rate significantly enhances grain-refining effects [12–14]. In order to illustrate this, samples were poured into steel cups of the reduced pressure testing machine (about 200 g) in air corresponding to a solidification rate of about

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<td>0.0014</td>
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Table 2. Chemical analysis of melt treated pure aluminum samples.
In the absence of a grain refiner, the macrostructure of pure aluminum is composed of three zones: (i) columnar near the bottom of the crucible due to the development of positive thermal gradient in the liquid metal leading to formation of grains with preferred orientation [14], (ii) fine equiaxed grained zone, and (iii) coarse grained zone where the grains are randomly oriented caused by fragmentation of the columnar grains driven by convection in the liquid metal, as illustrated in Figure 7 for samples solidified at 0.8°C/s. The equiaxed
crystals have a larger size than the chilled ones, and they are the result of volumetric solidification, which proceeds when the initially high thermal gradient is reduced as solidification progresses [15]. The refinement of the primary structure is a result of the creation of phases that act as substrates of heterogeneous nucleation for the primary aluminum phase. Therefore, active centers of aluminum heterogeneous nucleation are particles that have a high melting point and close crystallographic match with aluminum, e.g., TiC, TiN, TiB, TiB_2, AlB_2, Al_3Ti, and Zn_3Ti [16–20].

When an Al-5%Ti-1%B master alloy is added to pure aluminum, it decomposes into Al_3Ti and TiB_2 phase particles. Since the dissolution of Al_3Ti is rather fast in Al-Si alloys, it forms Ti(Si_{1−x},Si_x)^3 where x < 0.15 [21]. The latter tends to cover the TiB_2 particle surface acting as a heterogeneous nucleation substrate for the α-Al through the peritectic reaction. The addition of 600 ppm Ti in the form of Al-10%Ti to pure aluminum (Figure 7(b)) resulted in the complete elimination of the columnar grains with marked reduction in the grain size of the two equiaxed zones observed in Figure 7(a). Increasing the Ti concentration to 1000 ppm led

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Table 3. WDS analysis (at.%) of AlB_2 and AlB_12 phases observed in Figure 3(a).

Figure 4. Presence of B and Si in Al-7%Si treated with 800 ppm B. Note the geometric shape of the particles in zone A.
to formation of one zone characterized by its ultra-fine grain size (Figure 7(c)). Similar observations were made when the liquid aluminum was treated with Al-5%-Ti-1%-B master alloy (Figure 7(d) and (e)).

As previously mentioned, B has no grain-refining effect on ultra-pure aluminum. However, if its addition is associated with a relatively high cooling rate, i.e., 10°C/s as shown in Figure 8, it could result in some refining depending on B concentration as shown in Figure 8(d), where at least 1000 ppm B is used to obtain more or less same level of grain size produced by the addition of 30 ppm TiB₂ (Figure 7(e)). These experiments always produced a grain size gradient from the bottom of the crucible to the top surface. This is usually attributed to the fact that the crucible surface acts as a nucleation substrate for heterogeneous nucleation of α-Al. In addition, once solidification initiates on the relatively cold bottom of the crucible, sensible and latent heat are released to the crucible, reducing the temperature gradient and effectively slowing
Figure 6. (a) Decomposition of Al-5%Ti-1%B master alloy, (b) backscattered electron image and (c) element distribution of Ti-B interaction in 1050 aluminum treated with 240 ppm B. Note the attraction of Ti ring toward the cluster of B-rich particles.

Figure 7. Macrostructures of ultra-pure aluminum treated by two types of grain refiners: (a) Al pure (no treatment), (b) Al pure + Al₃Ti (600 ppm), (c) Al pure + Al₃Ti (1000 ppm), (d) Al pure + TiB₂ (15 ppm), (e) Al pure + TiB₂ (30 ppm). All micrographs have same magnification-solidification time was ~200 s. Note the heterogeneity in grain size on going from bottom to top.
the solidification rate. According to Chalmers [15], predendritic solid nuclei that formed on the mold surface during filling of the mold are driven away by circulation of the liquid metal. Depending on the degree of undercooling, some of these nuclei would survive in the liquid until the undercooling is extracted, resulting in equiaxed grains.

Considering the Al-7%Si alloy, its structure was not affected by the phenomenon mentioned above in the case of ultra-pure aluminum. Figure 9 shows the macrostructure of the alloy with different melt treatments consisting of fine equiaxed grains caused by the addition of a suitable dose of grain refiner associated with application of high solidification rate—note
the absence of columnar grains. The transition from columnar to equiaxed grains depends on several factors: (i) mold geometry, (ii) the amount of extracted heat, (iii) composition of used alloy, (iv) density of the used grain refiner, (v) fluidity of the liquid metal, (vi) interaction between grain refiner and alloy composition and (vii) sedimentation of the added grain refiner [22]. A high solidification rate, large amount of grain refiners as well as low temperature gradient would enhance fine equiaxed grains [23]. Al-7wt.% Si alloys with and without grain refiners solidified in diffusive conditions showed columnar growth in case of nonrefined alloy, and the existence of a columnar-to-equiaxed transition (CTE) in refined alloy. A sharp CTE is
observed when increasing the solidification rate and a progressive CTE when lowering the temperature gradient [24].

4. Concluding remarks

The results obtained show that when Al-4%B is added to ultra-pure aluminum, it decomposes into AlB$_{12}$ and AlB$_2$ phases that have no grain-refining effect in pure aluminum. When no grain refiner is added to pure aluminum, the microstructure is a mixture of columnar and equiaxed grains. The addition of 30 ppm Ti to pure aluminum is sufficient to promote equiaxed grains when the metal is solidified at high rate (~10°C/s). A similar effect can be achieved with the addition of excessive amounts of B, in the order of 1000 ppm. Increasing the Si content to 7% enhances the reduction in initial grain size of pure aluminum from 2800 to 1850 μm in Al-7%Si alloy. In the case of Al-7%Si, only 500 ppm B is sufficient to reduce the initial grain size by about 75%. Due to a reduced temperature gradient as the solidification front travels from the bottom of the crucible to the top, the grain size varies accordingly.

In commercial aluminum, the addition of B would react with traces of Ti leading to the formation of Al$_3$Ti and TiB$_2$ phases that are effective grain-refining agents. In the case of Al-7%Si, Ti reacts with Si forming (Al, Si)$_2$Ti phase which is a poor refining agent. This phenomenon is termed poisoning. No interaction was observed to take place between B and Si. In general, the grain size varied along the height of the crucible due to change in the liquid temperature.

Acknowledgements

The authors would like to thank Amal Samuel for enhancing the quality of the images and drawings presented in this article and Dr. E. Samuel for editing the manuscript.

Conflict of interest

There is no conflict of interest between the authors.

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References


